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(21) International Application Number: PCT/US96/19642 (22) International Filing Date: 12 December 1996 (12.12.96) (30) Priority Data: 08/577,355 22 December 1995 (22.12.95) US (71) Applicant: E.I. DU PONT DE NEMOURS AND COMPANY [US/US]; 1007 Market Street, Wilmington, DE 19898 (US). (72) Inventors: BREIKSS, Anne, Irise; 16 Cinnamon Drive, Hockessin, DE 19707-1349 (US). FOO, Thomas; Apart- ment 2-A-3, 402 Foulk Road, Wilmington, DE 19803-3805 (US). (74) Agent: DEITCH, Gerald, E.; E.I. du Pont de Nemours and Company, Legal Patent Records Center, 1007 Market Street, Wilmington, DE 19898 (US).		(81) Designated States: BR, CA, CN, JP, KR, SG, European patent (AT, BE, CH, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE). Published <i>With international search report.</i>
(54) Title: IMPROVED PROCESS FOR HYDROCYANATION OF DIOLEFINS AND ISOMERIZATION OF NONCONJUGATED 2- ALKYL-3-MONOALKENENITRILES (57) Abstract A process for the hydrocyanation of diolefins employing a catalyst comprising the combination of a zero-valent nickel compound and certain bidentate phosphorus compounds, and a process for isomerizing 2-alkyl-3-monoalkenenitrile to form linear nitriles by the use of such catalyst.		

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TITLE

Improved Process for Hydrocyanation of Diolefins
and Isomerization of Nonconjugated
2-Alkyl-3-Monoalkenenitriles

Field of the Invention

This invention generally relates to an improved liquid phase process useful in the hydrocyanation of diolefinic compounds to produce nonconjugated acyclic olefin nitriles and to the subsequent liquid phase process of isomerization of said nitriles to 3-and/or 4-monoalkene linear nitriles. The improvement resides in conducting the process in the presence of zero-valent nickel and a bidentate phosphorus compound.

Background of the Invention

Catalytic hydrocyanation systems, particularly pertaining to the hydrocyanation of olefins, are known in the art. For example, liquid phase systems useful for the hydrocyanation of butadiene to form pentenenitriles (PN) are known in the art. For example, Drinkard, U.S. 3,496,215, discloses the hydrocyanation of butadiene using monodentate nickel phosphite catalysts. As used in this patent, and as will be used herein, the term "pentenenitrile" is intended to mean cyanobutene. Likewise, "butenenitrile" means cyanopropene. Bidentate phosphite ligands complexed to zero-valent nickel and platinum are known to be useful in the liquid phase hydrocyanation of butadiene, as described by Baker et al. J. Chem. Soc., Chem. Commun., 1991, pages 803-804.

The pentenenitriles so formed are subjected to further hydrocyanation and/or isomerization to form adiponitrile (ADN), a commercially important material in the manufacture of nylon. For example, Drinkard, U.S. 3,536,748, discloses the liquid phase isomerization of 2-methyl-3-butenenitrile in the

presence of a zero-valent nickel complex, and Chia, U.S. 3,676,481, discloses an improvement additionally utilizing tri(hydrocarbyl)boron promoters.

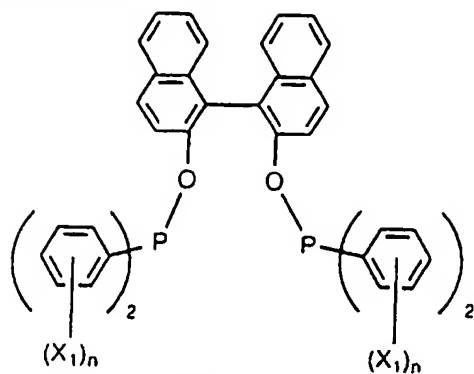
The hydrocyanation of activated olefins such as conjugated olefins (e.g., styrene) and strained olefins (e.g., norbornene) proceeds without the use of a Lewis Acid Promoter. For example, Casalnuovo, U.S. 5,175,335, discloses the use of chiral, nonracemic, bidentate phosphinite ligands for the enantioselective hydrocyanation of aromatic vinyl compounds. In contrast, the hydrocyanation of unactivated olefins such as 1-octene and 3-pentenitrile normally requires the use of a Lewis Acid Promoter. Teachings regarding the use of a promoter in the hydrocyanation reaction appear, for example, in U.S. 3,496,217.

Bidentate phosphinite ligands used in the present invention for the hydrocyanation of diolefins have been used for the hydrocyanation of monoolefins. Commonly assigned, copending Application Serial No. U.S. SN 08/408,250, filed March 22, 1995, disclose such bidentate phosphinite ligands used in combination with a Lewis Acid Promoter to hydrocyanate monoolefins.

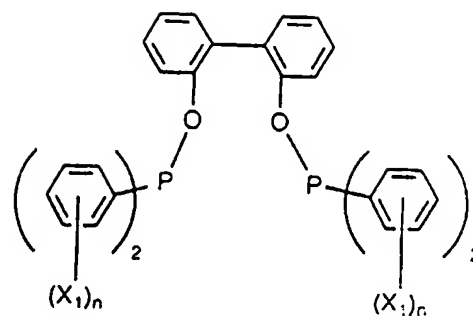
The present invention provides for an improved process for the hydrocyanation of diolefinic compounds, such as butadiene, and subsequent isomerization of monoolefins without the need for Lewis Acid Promoters utilizing zero-valent nickel and a bidentate phosphinite ligand. Other objects and advantages of the present invention will become apparent to those skilled in the art upon reference to the detailed description of the invention which hereinafter follows.

Summary of the Invention

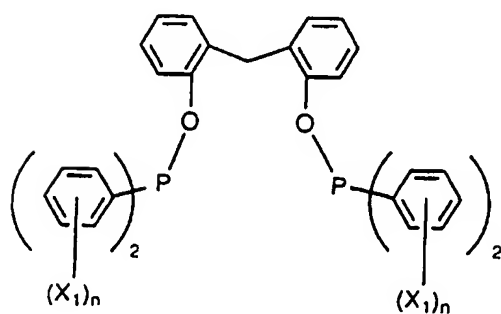
The present invention provides an improved process for the liquid phase hydrocyanation of diolefinic compounds and subsequent isomerization of the resulting nonconjugated acyclic nitriles comprising, reacting an aliphatic diolefinic compound, preferably butadiene, with a source of HCN, wherein the improvement comprises conducting the hydrocyanation and subsequent isomerization in the presence of a catalyst composition comprising zero-valent nickel and at least one bidentate phosphorus ligand selected from the group consisting of compounds represented by Formulas I through XIII:



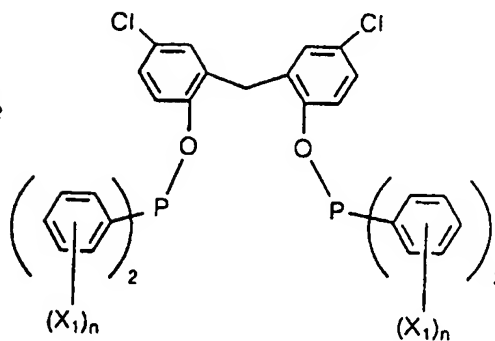
Formula I



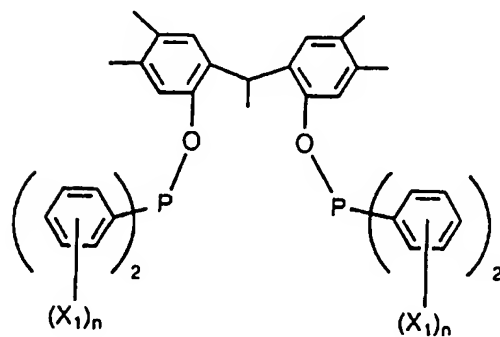
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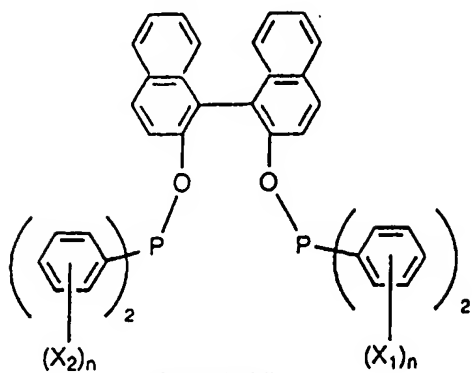
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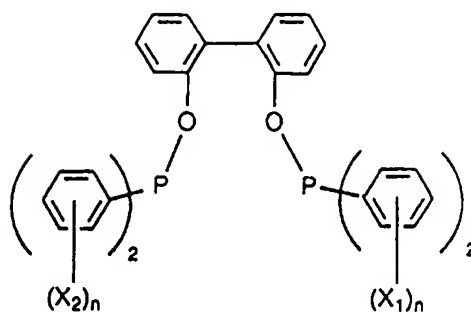
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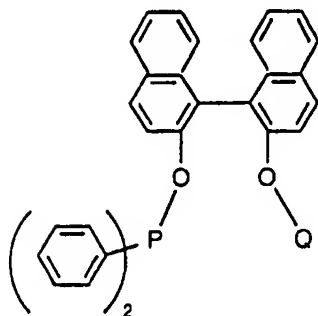
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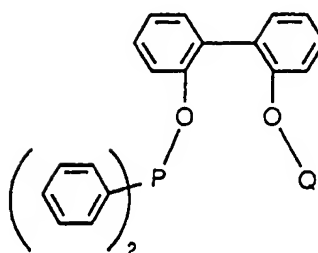
Formula VI



Formula VII

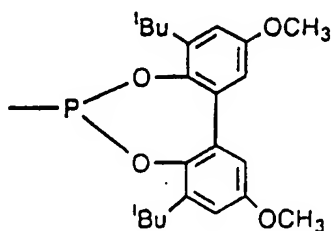
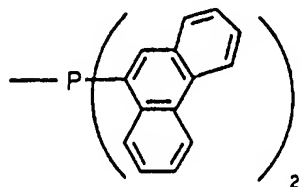
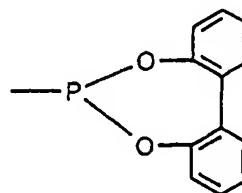
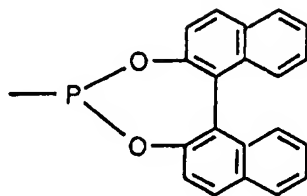


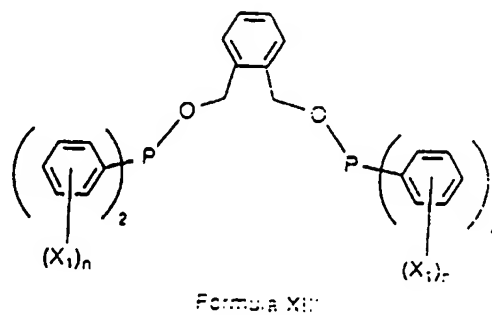
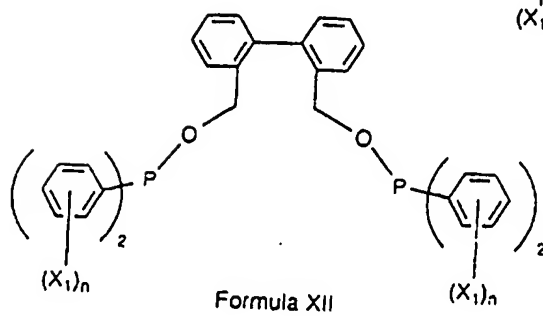
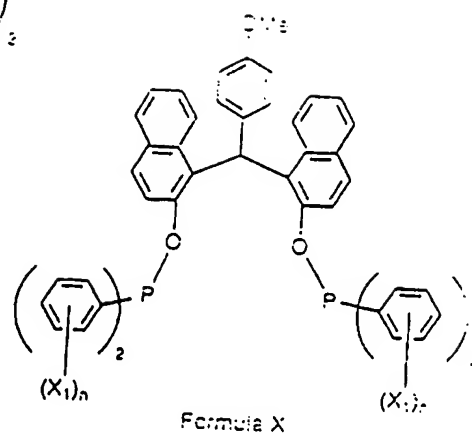
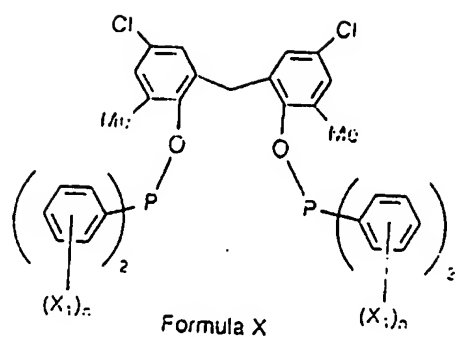
Formula VIII



Formula IX

where Q =





where X_1 is meta -Cl, para -Cl, meta -CF₃, para -CF₃, meta -F, para -F, meta -CN, para -CN, meta -CH₃, or para -CH₃; X_2 is methyl or alkoxy having 1 to 3 carbon atoms; n is zero, 1, or 2.

5 The reaction is most conveniently performed continuously from hydrocyanation of the starting diolefin to the final 3- and/or 4-monoalkene linear nitriles. However, the process can be conducted stepwise, i.e., the nonconjugated nitriles resulting
10 from the hydrocyanation can be isolated per se, prior to isomerization. Furthermore, nonconjugated acyclic olefin nitriles prepared by any method can be used as starting materials for the isomerization in accordance with this invention.

15

Detailed Description of the Invention

The catalyst precursor compositions useful for the practice of this invention are comprised of a bidentate phosphorus compound and zero-valent nickel.
20 The catalyst composition is referred to as a "precursor" only to indicate, in all likelihood during the hydrocyanation reaction, the structure of the active catalyst composition may, in fact, be complexed to an olefin.

25 The bidentate phosphorus compounds may be prepared by a variety of methods known in the art. The symmetrical diphosphetes are prepared as follows. The diarylchlorophosphine is added to a toluene solution of a diol and triethylamine. The reaction
30 mixture is allowed to stir at room temperature, then filtered to remove triethylamine hydrochloride. The product is isolated by removing the solvent under reduced pressure.

The unsymmetrical diphosphetes are prepared
35 in a similar manner. The first diarylchlorophosphine (preferably the more sterically hindered one) is added to a toluene solution of a diol and triethylamine. Once the reaction is complete, the second

diarylchlorophosphine is added. Triethylamine hydrochloride is filtered off and the solvent removed under reduced pressure to give the product.

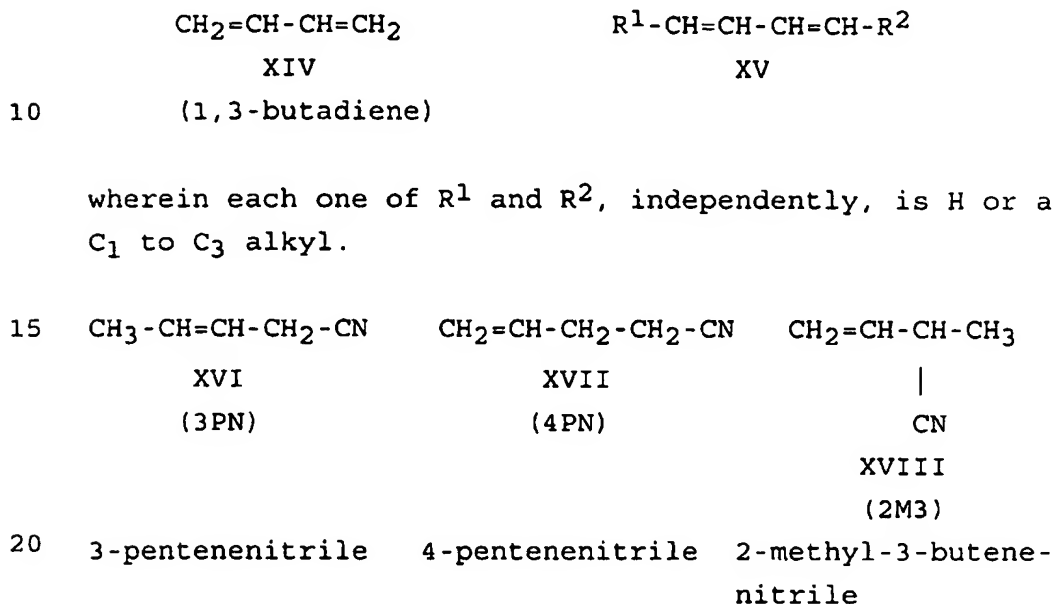
The zero-valent nickel can be prepared or
5 generated according to techniques known in the art (U.S. 3,496,217; 3,631,191; 3,846,461; 3,847,959; and 3,903,120, which are incorporated by reference). Zero-valent nickel compounds that contain ligands, which can be displaced by the organophosphorus ligand, are a
10 preferred source of zero-valent nickel. Two such preferred zero-valent nickel compounds are $\text{Ni}(\text{COD})_2$ (COD is 1,5-cyclooctadiene) and $\text{Ni}(\text{P}(\text{O}-\text{o}-\text{C}_6\text{H}_4\text{CH}_3)_3)_2(\text{C}_2\text{H}_4)$, both of which are known in the art. Alternatively, divalent nickel compounds may
15 be combined with a reducing agent and are then able to serve as suitable sources of zero-valent nickel in the reaction. Suitable divalent nickel compounds include compounds of the formula NiY_2 where Y is halide, carboxylate, or acetylacetonate. Suitable reducing
20 agents include metal borohydrides, metal aluminum hydrides, metal alkyls, Zn, Fe, Al, Na, or H_2 . Elemental nickel, preferably nickel powder, when combined with a halogenated catalyst, as described in U.S. 3,903,120, is also a suitable source of zero-
25 valent nickel.

The actual catalyst is a complex of zero-valent nickel with the bidentate ligand, which is formed when those two materials are combined. An effective catalyst requires at least two moles of P
30 atoms for one gram-atom of zero-valent nickel.

The diolefinic compound reactants used in this invention include primarily conjugated diolefins containing from 4 to 10 carbon atoms; for example, 1,3-butadiene, and cis- and trans-2,4-hexadienes, but also
35 includes cis- and trans-1,3-pentadienes and 1,5-cyclooctadiene. Butadiene is especially preferred by reason of its commercial importance in the production of adiponitrile. Other suitable diolefinic compounds

include diolefinic compounds substituted with groups which do not deactivate the catalyst.

The following Formulas XIV and XV illustrate suitable representative starting diolefinic compounds; and Formulas XVI, XVII, and XVIII represent the products obtained from 1,3-butadiene and HCN;



It will be recognized that compound XIV 1,3-butadiene is a special case of Formula XV, where each one of R^1 and R^2 is hydrogen.

In the practice of the hydrocyanation of the diolefin in accordance with the present invention, the following description applies.

The hydrocyanation reaction can be carried out with or without a solvent. Any solvent should be a liquid at the reaction temperature and inert towards the unsaturated compound and the catalyst. Generally, such solvents are hydrocarbons such as benzene, xylene, or nitriles such as acetonitrile, benzonitrile, or adiponitrile.

The exact temperature used is dependent, to a certain extent, on the particular catalyst being used, the particular unsaturated compound being used and the

desired rate. Generally, temperatures of from -25°C to 200°C can be used with from 0°C to 150°C, being the preferred range.

The reaction may be carried out by charging a reactor with all of the reactants, or preferably, the reactor is charged with the catalyst or catalyst components, the unsaturated compound and any solvent used, and the hydrogen cyanide gas is swept over the surface of the reaction mixture or bubbled through said reaction mixture. If desired, when using a gaseous unsaturated organic compound, the hydrogen cyanide and the unsaturated organic compound may be fed together into the reaction medium. The molar ratio of unsaturated compound to catalyst generally is varied from about 10:1 to 100,000:1, preferably 100:1 to 5,000:1, unsaturated compound to catalyst for a batch operation. In a continuous operation, such as when using a fixed bed-catalyst type of operation, a higher proportion of catalyst may be used such as 5:1 to 100,000:1, preferably 100:1 to 5,000:1, unsaturated compound to catalyst.

Preferably, the reaction mixture is agitated, such as by stirring or shaking.

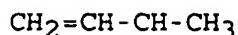
The cyanated product can be recovered by conventional techniques such as crystallization of the product from solution or by distillation.

One can either isolate the 2-alkyl-3-monoalkenenitriles produced by the hydrocyanation of the diolefin or proceed continuously with the isomerization under similar reaction conditions.

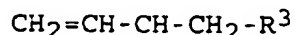
The 2-alkyl-3-monoalkenenitrile used as the starting materials in the isomerization of this invention can result from the hydrocyanation of diolefin described above or can come from any other available source. The olefinic double bond in the 2-alkyl-3-monoalkenenitrile used as the starting materials in the isomerization of this invention cannot be conjugated to the triple bond of the cyano group.

Suitable starting 2-alkyl-3-monoalkenenitriles can also carry groups which do not attack the catalyst, for example, another cyano group. Preferably, the starting 2-alkyl-3-monoalkenenitriles contain from 5 to 8 carbon atoms, excluding any additional substitution. 2-methyl-3-butenenitrile is especially important in the production of adiponitrile. Other representative nitriles include 2-ethyl-3-butenenitrile and 2-propyl-3-butenenitrile.

The following Formulas XIX and XX illustrate suitable representative starting 2-alkyl-3-monoalkenenitriles. When the starting nitrile is 2-methyl-3-butenenitrile, the isomerization products are those shown in Formulas XXI and XXII;



XIX
(2M3)

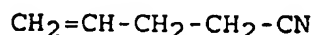


XX

wherein R^3 is H or a C_1 to C_3 alkyl.



XXI
(3PN)



XXII
(4PN)

It will be recognized that Formula XIX is a special case of Formula XX, where R^3 is hydrogen.

The isomerization process of this invention can be carried out, for example, at atmospheric pressure and at any temperature in the range of 10-200°C, preferably in the range of 60-150°C. The pressure is not critical, however, and can be above or below atmospheric pressure, if desired. Any of the conventional batch or continuous flow procedures may be used either in the liquid phase or in the vapor phase (with respect to the relatively volatile 2-methyl-3-

butenenitrile reactant and linear pentenenitrile products). The reactor may be of any mechanically and chemically resistant material and is usually of glass or an inert metal or alloy, e.g., nickel, copper, silver, gold, platinum, stainless steel, Monel®, Hastelloy®, etc.

The process is usually carried out "neat," i.e., without an added diluent or solvent. Any solvent or diluent that is nondestructive of the catalyst can be used, however. Suitable solvents include aliphatic or aromatic hydrocarbons (hexane, cyclohexane, benzene), ethers (diethyl ether, tetrahydrofuran, dioxane, glycol dimethyl ether, anisole), esters (ethyl acetate, methyl benzoate), nitriles (acetonitrile, benzonitrile), etc.

A nonoxidizing environment is desirable in order to retard oxidative deactivation of the catalyst. Accordingly, an inert atmosphere, e.g., nitrogen, is normally and preferably used, although air may be used, if desired, at the expense of loss of a proportion of the catalyst through oxidation.

When the process is a typical batch operation in the liquid phase with or without a solvent, the catalytic nickel complex is soluble to some extent at temperatures within the operating range and is usually completely soluble at the most preferred operating temperature. However, the nickel complex is essentially nonvolatile, whereas the 2-methyl-3-butenitrile reactant, and the linear pentenenitrile products are relatively volatile. Accordingly, in a continuous flow procedure, the catalyst may be a component of the flowing system in a completely liquid-phase operation, it may be in a mobile nonflowing liquid state in a semi-vapor-phase operation, or it may be in a fixed-bed state (usually on a solid support) in a conventional flowing vapor-phase operation.

The time element in the process is not critical and may generally be governed by practical

considerations. The time required for a practical level of conversion of 2-methyl-3-butenitrile to linear pentenenitriles is dependent upon the temperature of the reaction, i.e., operation at lower temperature generally requires a longer time than operation at a higher temperature. A practical reaction time can be in the range of a few seconds to many hours, depending on the particular conditions and method of operation.

10 The molar ratio of 2-methyl-3-butenitrile to catalyst is generally greater than 1:1, usually in the range from about 5:1 to 20,000:1, preferably 100:1 to 5,000:1, for a batch or continuous operation.

In a preferred process, the bidentate
15 phosphorus compound is of Formula I, where each X is CF_3 , Cl, or F and $n = 0, 1$, or 2.

EXAMPLES

The invention will now be illustrated by the
20 following non-limiting examples of certain preferred embodiments thereof, wherein all parts, proportions, and percentages are by weight, unless otherwise indicated.

25 Example 1

A 25 wt% solution of 1,3-butadiene was made by vacuum transfer of 9.11 g of 1,3-butadiene into 27.33 g of butyronitrile. A 25 wt% solution of HCN was made by the addition of 2.506 g of HCN to 7.503 g of propionitrile. The catalyst solution was prepared by the addition of 0.055 g of $\text{Ni}(\text{COD})_2$ and 0.474 g of the ligand having the structure labeled "Example 1" below to 9.47 g of propionitrile. With these solutions, the following reaction mixtures were prepared in 2-ml GC
35 vials equipped with micro-stirbars:

	<u>Sample 1</u>	<u>Sample 2</u>	<u>Sample 3</u>
Butadiene Solution	0.206 g	0.202 g	0.205 g
HCN Solution	0.080 g	0.081 g	0.085 g
Catalyst Solution	0.080 g	0.077 g	0.077 g

The GC vials were crimp-sealed with caps fitted with appropriately sized sheets of Nordel rubber to aid in containing the reaction mixtures. The vials were placed in a hot-block-stirrer set at 80°C. Sample 1 was removed after 1.0 h of reaction time. Sample 2 was removed after 2.0 h of reaction time. Sample 3 was removed after 3.0 h of reaction time. The reaction was quenched in each case by diluting the reaction mixture with diglyme as a GC solvent for product analysis. The propionitrile in the reaction mixture was used as an internal standard in the GC product analysis. The results of the analyses are shown in Table 1.

TABLE 1
BUTADIENE HYDROCYANATION

	<u>EXAMPLE</u>	<u>TIME</u>	<u>% 3PN</u>	<u>% 2M3</u>	<u>TOTAL PN</u>
	1	1 hr.	34.9%	70.8%	105.6%
		2 hr.	34.8%	68.8%	103.5%
20		3 hr.	35.6%	70.9%	106.4%
	2	1 hr.	34.3%	71.0%	105.3%
		2 hr.	34.9%	68.4%	103.3%
		3 hr.	35.6%	65.4%	101.0%
25	3	1 hr.	33.4%	66.8%	100.2%
		2 hr.	32.3%	64.2%	96.5%
		3 hr.	31.2%	62.8%	94.0%
30	4	1 hr.	36.1%	64.3%	100.4%
		2 hr.	31.9%	59.1%	91.0%
		3 hr.	31.0%	61.7%	92.7%
	5	1 hr.	34.0%	57.3%	91.3%
35		2 hr.	33.5%	56.9%	90.4%
		3 hr.	34.0%	57.1%	91.1%
	6	1 hr.	18.4%	55.1%	73.5%
		2 hr.	21.3%	63.0%	84.3%
40		3 hr.	21.9%	65.0%	87.0%
	7	1 hr.	26.3%	69.3%	95.6%
		2 hr.	27.3%	68.2%	95.5%
		3 hr.	21.7%	59.6%	81.3%

TABLE 1
BUTADIENE HYDROCYANATION CONT'D

	<u>EXAMPLE</u>	<u>TIME</u>	<u>% 3PN</u>	<u>% 2M3</u>	<u>TOTAL PN</u>
5	8	30	33.4%	69.7%	103.0%
		1 hr.	32.0%	62.9%	94.9%
		2 hr.	28.8%	51.9%	80.7%
	9	1 hr.	7.6%	49.9%	57.5%
10		2 hr.	9.6%	52.9%	62.5%
		3 hr.	10.3%	55.0%	65.3%
	10	1 hr.	6.0%	14.0%	20.0%
		2 hr.	10.6%	24.2%	34.8%
15		3 hr.	14.0%	31.8%	45.8%
	11	1 hr.	2.5%	8.8%	11.2%
		2 hr.	8.0%	20.1%	28.1%
		3 hr.	12.9%	30.2%	43.1%
20	12	1 hr.	7.9%	25.8%	33.7%
		2 hr.	8.8%	27.9%	36.7%
		3 hr.	8.4%	27.3%	35.7%
25	13	1 hr.	11.7%	21.8%	33.5%
		2 hr.	11.1%	20.7%	31.9%
		3 hr.	11.8%	22.0%	33.7%
	14	1 hr.	8.3%	23.7%	32.0%
30		2 hr.	9.2%	23.4%	32.6%
		3 hr.	9.4%	22.4%	31.8%
	15	1 hr.	8.0%	24.9%	32.9%
		3 hr.	7.7%	24.2%	31.9%
35		4 hr.	7.4%	23.5%	30.9%

In the examples as shown in Table 1, the butadiene hydrocyanation experiments were performed as follows.

40 Example 2 through 15 were run in a manner similar to Example 1, except that the solvent for the components might have differed. All solvents for these examples were chosen from propionitrile, butyronitrile, or hexanenitrile.

45 In the examples as shown in Table 2, the 2M3 isomerization experiments were performed as follows. A sample containing 81 to 83% of 2-methyl-3-butenenitrile

(2M3), with other pentenenitrile isomers, was the starting material.

Example 16

5 The catalyst solution was prepared by the addition of 0.055 of Ni(COD)₂ and 0.439 g of the ligand having the structure labeled "Example 16" below to 9.51 g of propionitrile. A sample of 2M3BN was distilled under nitrogen onto 100 parts per million of
10 2,6-di-tert-butyl-4-methylphenol. GC analysis of this sample showed it to be 81% 2M3BN. With these mixtures, the following reaction mixtures were prepared in 2-ml GC vials equipped with micro-stirbars:

	<u>Sample 0</u>	<u>Sample 1</u>	<u>Sample 2</u>
15 Catalyst Solution	0.085 g	0.112 g	0.101 g
2M3BN	0.084 g	0.111 g	0.101 g

20 The GC vials were crimp-sealed with caps fitted with appropriately sized sheets of Nordel rubber to aid in containing the reaction mixtures. Samples 1-2 were placed in a hot-block-stirrer set at 125°C. Sample 0 represented the t = 0. Sample 1 was removed after 1.0 h of reaction time. Sample 2 was removed after 2.0 h of reaction time. The reaction was
25 quenched in each case by diluting the reaction mixture with diglyme as a GC solvent for product analysis. The propionitrile in the reaction mixture was used as an internal standard in the GC product analysis. The results of the analyses are shown in Table 2.

TABLE 2

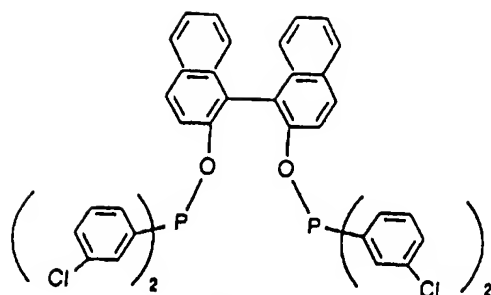
2-METHYL-3-BUTENENITRILE ISOMERIZATION

	<u>EXAMPLE</u>	<u>TIME</u>	<u>% 2M3</u>	<u>% 3PN</u>
5	16	0 hr.	106.5%	1.7%
		1 hr.	41.7%	66.6%
		2 hr.	38.0%	71.5%
	17	0 hr.	101.2%	1.9%
10		1 hr.	44.5%	59.4%
		2 hr.	43.6%	68.5%
		3 hr.		70.0%
	18	0 hr.	98.6%	1.4%
15		1 hr.	43.9%	50.8%
		2 hr.	39.9%	55.4%
		3 hr.	41.5%	56.2%
	19	0 hr.	89.5%	1.0%
20		1 hr.	64.3%	34.7%
		2 hr.	68.2%	37.6%
		3 hr.	56.4%	33.8%
	20	0 hr.	100.1%	1.2%
25		1 hr.	85.0%	13.1%
		2 hr.	76.9%	16.6%
		3 hr.	63.1%	30.7%

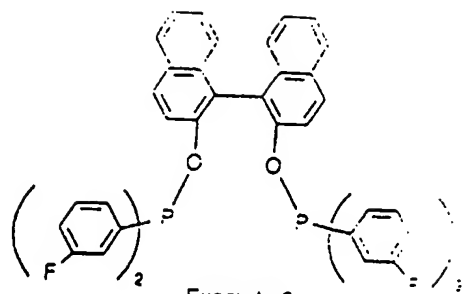
Examples 17-20 were run in a manner similar
30 to Example 16 using the ligands designated below.

The structures of bidentate phosphorous
compounds used in the examples are shown below:

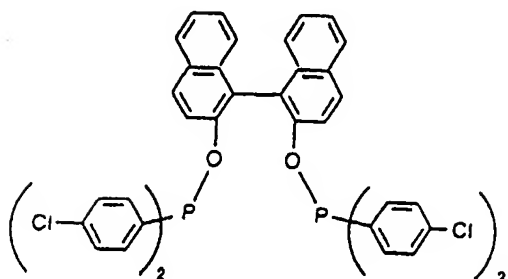
Structures for Examples in Table 1



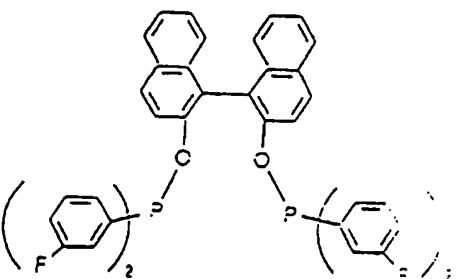
Example 1



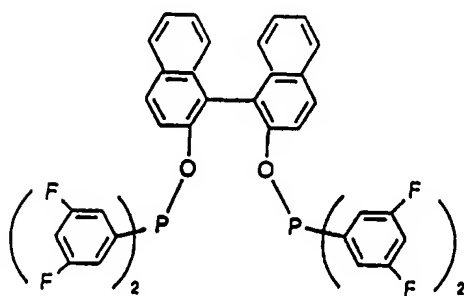
Example 2



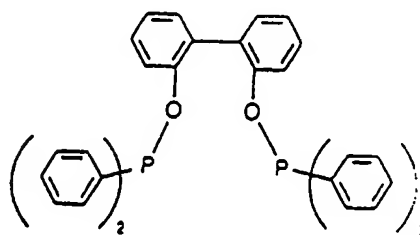
Example 3



Example 4



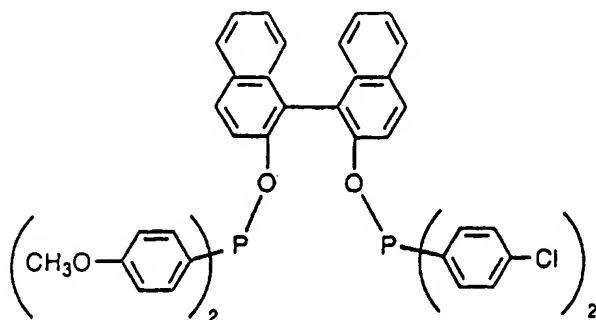
Example 5



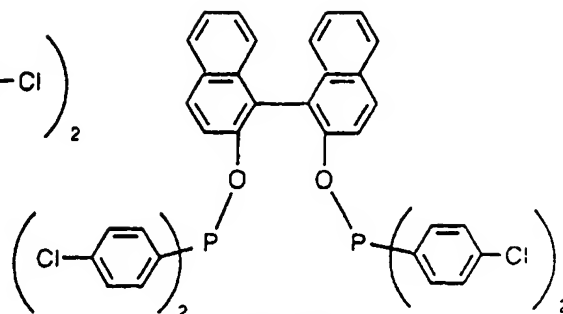
Example 6

Structures for Examples in Table 1

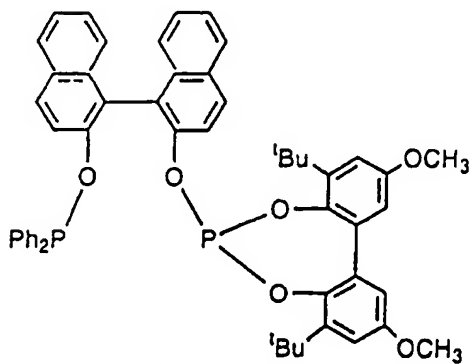
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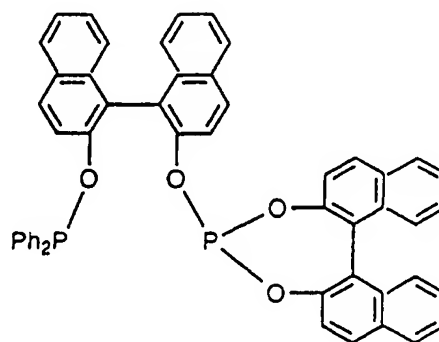
Example 7



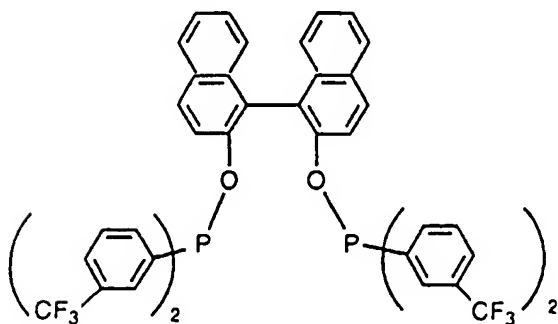
Example 8



Example 9



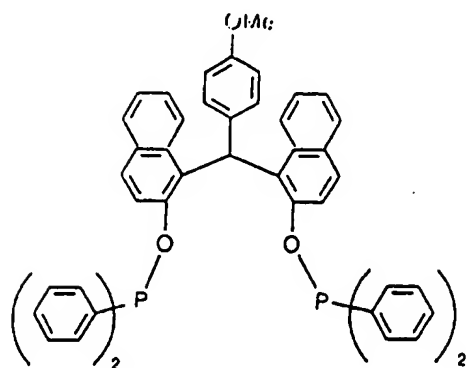
Example 10



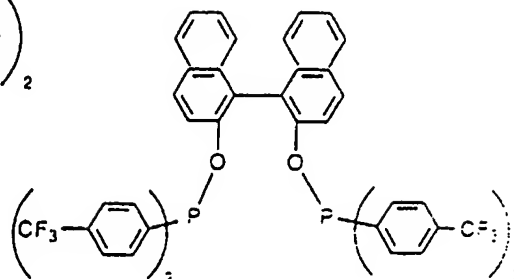
Example 11

Structures for Examples in Table 1

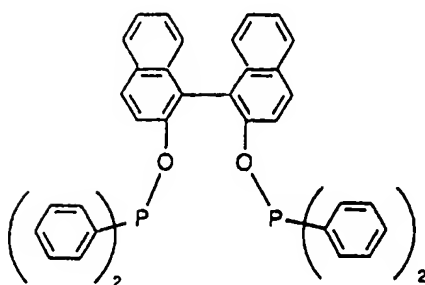
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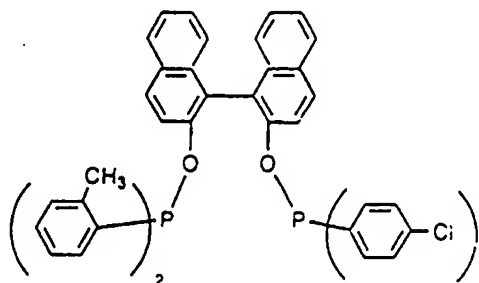
Example 12



Example 13

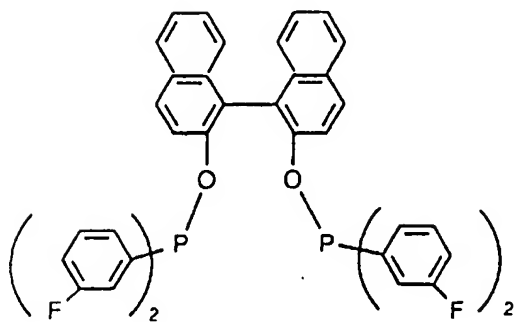


Example 14

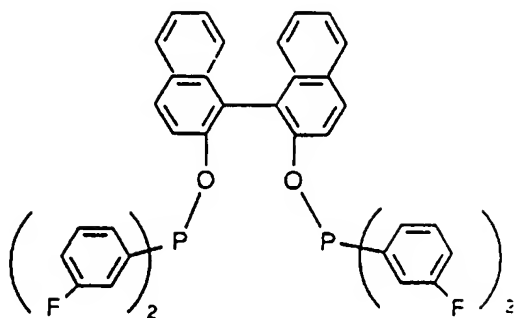


Example 15

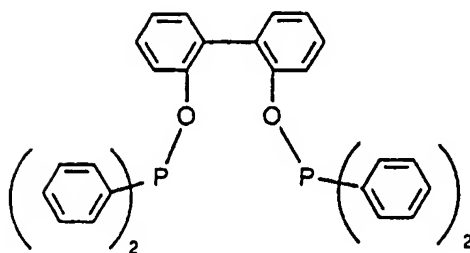
Structures for Examples in Table 2



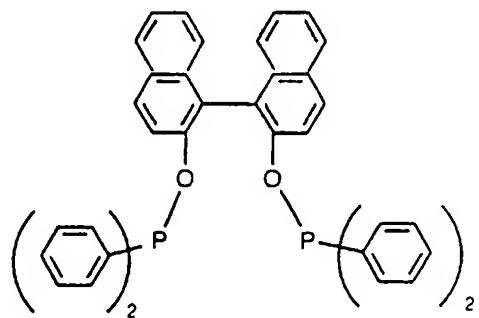
Example 16



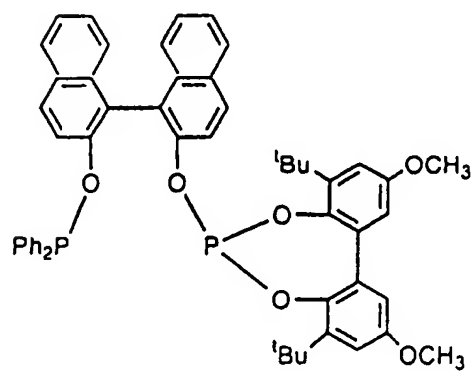
Example 17



Example 18



Example 19



Example 20

What is claimed is:

1. A process for the liquid phase hydrocyanation of an aliphatic diolefin having 4 to 10 carbon atoms which comprises reacting said diolefin with HCN at a temperature in the range of -25 to 200°C in the presence of a catalyst comprising a zero-valent nickel compound and at least one bidentate phosphorus compound selected from compounds having the Formulas I through XIII.
2. The process of claim 1 in which the diolefin is selected from the group consisting of 1,3-butadiene, cis and trans-2,4-hexadienes, cis and trans-1,3-pentadienes, and 1,5-cyclooctadiene.
3. The process of claim 1 in which the diolefin is a conjugated acyclic aliphatic diolefin containing 4 to 10 carbon atoms.
4. The process of claim 1 in which the molar ratio of HCN to catalyst is between 100:1 and 5000:1.
5. The process of claim 1 in which the bidentate phosphorus compound is selected from Formula I and Formula II.
6. The process of claim 5 in which the bidentate phosphorus compound is from Formula I and each X_1 is CF_3 , Cl, or F and $n = 0, 1, \text{ or } 2$.
7. A process which comprises the liquid phase hydrocyanation of a aliphatic diolefin having 4 to 10 carbon atoms to form a mixture containing nonconjugated acyclic nitriles including some 2-alkyl-3-monoalkenenitrile by reacting said diolefin with HCN at a temperature in the range of -25 to 200°C in the presence of a catalyst comprising a zero-valent nickel compound and at least one bidentate phosphorus compound selected from compounds having the Formulas I through XIII, and then isomerizing the 2-alkyl-3-monoalkenenitrile contained in the mixture to form 3- and/or 4-monoalkene linear nitriles by reacting said mixture in the presence of the catalyst comprising the

nickel compound and the bidentate phosphorus compounds at a temperature in the range of 10 to 200°C.

8. A process which comprises isomerizing the 2-alkyl-3-monoalkenenitrile contained in a mixture
5 of nonconjugated nitriles to form 3- and/or 4-
monoalkene linear nitriles by reacting said mixture in the presence of the catalyst comprising a zero-valent nickel compound and at least one bidentate phosphorus compound selected from compounds having the Formulas I
10 through XIII at a temperature in the range of 10 to 200°C.

9. The process of claim 1 in which the aliphatic diolefin is butadiene.

INTERNATIONAL SEARCH REPORT

onal Application No

PCT/US 96/19642

A. CLASSIFICATION OF SUBJECT MATTER
IPC 6 C07C253/10 C07C255/07

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 6 C07C

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	JOURNAL OF THE CHEMICAL SOCIETY, CHEMICAL COMMUNICATIONS, 1991, LETCHWORTH GB, pages 803-4, XP002002591 M. J. BAKER ET AL.: "Chelating diphosphite complexes of nickel(0) and platinum(0): their remarkable stability and hydrocyanation activity" cited in the application see the whole document ---	1-4,9
A	US 3 686 264 A (ALBANESE PIETRO ET AL) 22 August 1972 see the whole document ---	1-4,9
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Date of the actual completion of the international search

20 February 1997

Date of mailing of the international search report

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INTERNATIONAL SEARCH REPORT

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A	US 5 175 335 A (CASALNUOVO ALBERT L ET AL) 29 December 1992 cited in the application see column 19; example 38 ---	1
A	EP 0 647 619 A (RHONE POULENC CHIMIE) 12 April 1995 see page 4, line 31 - line 41 ---	7
A	WO 95 30680 A (DSM NV ;KAMER PAULUS CLEMENS JOZEF (NL); KRANENBURG MIRKO (NL); LE) 16 November 1995 see page 15, line 4 - page 16, line 29; ligand 20 - 31 -----	1

INTERNATIONAL SEARCH REPORT

information on patent family members

International Application No

PCT/US 96/19642

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		NL-A- 7002580	07-09-70

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